

Review Exercise 4 - Naming Binary Compounds



Here is a mix of all the kinds of binary compounds you have reviewed. Complete this exercise before you check your answers. Then make corrections as necessary.

1. Fill in the chart with the required information. The first one is done for you.

Atoms Combined	Formula of ions involved (if ionic)	Chemical Formula	Name of Compound
1. Ca and Cl	Ca^{2+} and Cl^-	CaCl_2	calcium chloride
2. K and Br	.	.	.
3. Mg and O	.	.	.
4. Al and O	.	.	.
5.	.	.	nitrogen trihydride
6. Ga and S	.	.	.
7.	.	$\text{HBr}_{(\text{aq})}$.
8. Fe^{3+} and O	.	.	.
9. Sn^{2+} and F	.	.	.
10. Cu^+ and N	.	.	.

2. Fill in the chart with the required information. If the formula is given, provide the name, and if the name is given, provide the correct formula.

Name of Compound	Chemical Formula	Chemical Formula	Name of Compound
1. nitrogen trihydride	.	16. CaBr_2	.
2. aluminum sulfide	.	17. BaO_2	.
3. hydrofluoric acid	.	18. CH_4	.
4. hydrogen sulfide gas	.	19. $\text{HI}_{(\text{aq})}$.
5. tin(IV) chloride	.	20. P_4	.
6. boron tetrahydride	.	21. Cr_2S_3	.
7. hydrogen peroxide	.	22. Hg_2O	.
8. lithium bromide	.	23. SnBr_2	.
9. calcium fluoride	.	24. AlCl_3	.
10. copper(II) nitride	.	25. K_2Te	.
11. antimony(V) oxide	.	26. $\text{H}_2\text{S}_{(\text{g})}$.
12. hydroselenic acid	.	27. N_2O_4	.
13. oxygen gas	.	28. CH_4	.
14. carbon tetrachloride	.	29. NaI	.
15. lead(IV) sulfide	.	30. Sb_2O_3	.